

CANDLE COMBUSTION

Student Class Date

MATERIALS

Two wax candles, two plates, lighter, small jar, big jar

PROCEDURE

Light the candles and put them on the plates.

Cover the candles with the jars and wait.

OBSERVATIONS

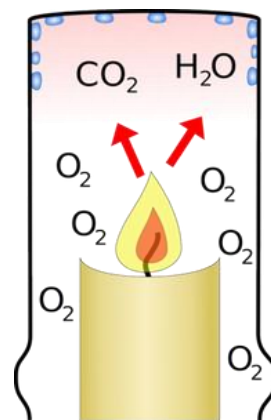
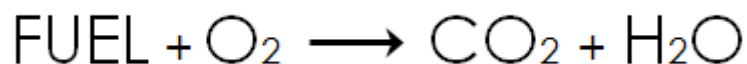
Tick the correct answers (more than one is correct):

- the flame in the smaller jar goes out after the bigger one.
- the flame in the smaller jar goes out before the bigger one.
- I observe a few condensation on the internal surface.
- I observe a lot of condensation on the external surface.
- I observe no condensation on the internal surface.

CONCLUSIONS

When a candle burns, the reactants are fuel (the candlewick and wax) and oxygen (in the air). The products are carbon dioxide gas and water vapor.

This is a **combustion reaction**. The general equation for this type of reaction is:



Why does the flame go out when we put a jar over the candle?

Tick the correct answer:

- the amount of hydrogen in the jar is not enough to allow a combustion.
- the water vapor blows out the candle.
- the oxygen runs out during the reaction and because of that, the flame goes out.
- the flame won't stop to burn.

Draw a picture!